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Chapter 22. $\rho = p/RT = (1.2)(1.01 \times 10^5$

$)/(287)(300) \rho = 1.41 \text{ kg/m}^3$

$v = 1/\rho = 1/1.41 = 0.71 \text{ m}^3/\text{kg}$

$2.23 \text{ kT} = (1.38 \times 10^{-23})$

$(500) = 1.035 \times 10^{-20} \text{ J}$

22 One kg-mole, which has a mass of 4 kg, has 6.02×10^{26} atoms.

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